**Chemistry Unit C1: Structures, Trends, Chemical Reactions, Quantitative Chemistry and Analysis**

**C1.1 Atomic structure**

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| **Content - CCEA Double Award Chemistry 1 – Fort Hill Integrated College** | Got it | Nearly | Haven’t a clue |
| **C1.1 Atomic structure** |
| Can you describe the structure of an atom as a central positively charged nucleus containing protons and neutrons (most of the mass) surrounded by orbiting electrons in shells; |  |  |  |
| Can you state the relative charges and approximate relative masses of protons, neutrons and electrons; |  |  |  |
| Can you define atomic number as the number of protons in an atom; |  |  |  |
| Can you define mass number as the total number of protons and neutrons in an atom; |  |  |  |
| Can you describe and explain that an atom as a whole has no electrical charge because the number of protons is equal to the number of electrons;  |  |  |  |
| Can you calculate the number of protons, neutrons and electrons in an atom or an ion and deduce the charge on an ion or determine the number of subatomic particles given the charge. |  |  |  |
| Can you write and draw the electronic configuration (structure) of atoms and ions with atomic number 1–20; |  |  |  |
| Can you define isotopes as atoms of an element with the same atomic number but a different mass number, indicating a different number of neutrons; |  |  |  |
| Can you interpret data on the number of protons, neutrons and electrons to identify isotopes of an element; |  |  |  |
| **Can you calculate the relative atomic mass of elements from the mass number and abundances of its isotopes**; |  |  |  |
| Can you recall that a compound is two or more elements chemically combined. |  |  |  |