**Chemistry Unit C1: Structures, Trends, Chemical Reactions, Quantitative Chemistry and Analysis**

**C1.1 Atomic structure**

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| --- | --- | --- | --- |
| **Content - CCEA Double Award Chemistry 1 – Fort Hill Integrated College** | Got it | Nearly | Haven’t a clue |
| **C1.1 Atomic structure** |
| Can you describe the structure of an atom as a central positively charged nucleus containing protons and neutrons (most of the mass) surrounded by orbiting electrons in shells; |  |  |  |
| Can you state the relative charges and approximate relative masses of protons, neutrons and electrons; |  |  |  |
| Can you define atomic number as the number of protons in an atom; |  |  |  |
| Can you define mass number as the total number of protons and neutrons in an atom; |  |  |  |
| Can you describe and explain that an atom as a whole has no electrical charge because the number of protons is equal to the number of electrons;  |  |  |  |
| Can you calculate the number of protons, neutrons and electrons in an atom or an ion and deduce the charge on an ion or determine the number of subatomic particles given the charge. |  |  |  |
| Can you write and draw the electronic configuration (structure) of atoms and ions with atomic number 1–20; |  |  |  |
| Can you define isotopes as atoms of an element with the same atomic number but a different mass number, indicating a different number of neutrons; |  |  |  |
| Can you interpret data on the number of protons, neutrons and electrons to identify isotopes of an element; |  |  |  |
| **Can you calculate the relative atomic mass of elements from the mass number and abundances of its isotopes**; |  |  |  |
| Can you recall that a compound is two or more elements chemically combined. |  |  |  |

Atoms, elements and compounds

All materials are made up of tiny particles called \_\_\_\_\_\_. All materials can be sorted out into 3 different groups depending on how their atoms are arranged;

* + - 1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
			2. \_\_\_\_\_\_\_\_\_\_\_
			3. \_\_\_\_\_\_\_\_\_\_\_

**Elements** are made of only **\_\_\_\_** type of atom.

**Compounds** consist of two or more different atoms **\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_** together.

**Mixtures** are made of different types of atom but these are **\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_** and are not chemically bonded.

|  |  |  |
| --- | --- | --- |
|  |  |  |
| Iron, Fe | Water, H2O | Air, N2 + O2 |

|  |  |  |
| --- | --- | --- |
|  |  |  |
| Sulphur, S | Carbon dioxide CO2 | Sulphur, S and Iron, Fe |

We use letters to stand for elements. We call these \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. For example, Na stands for one atom of \_\_\_\_\_\_\_\_\_\_\_\_\_\_ and O stands for one atom of \_\_\_\_\_\_\_\_\_\_\_\_\_\_. The periodic table shows all of the elements. Each column contains elements with similar \_\_\_\_\_\_\_\_\_\_\_\_\_\_. We call each column a \_\_\_\_\_\_\_\_\_\_\_\_\_.



**What’s in a compound?**

When atoms of different elements join together we get a substance called a \_\_\_\_\_\_\_\_\_\_\_\_. Most substances are compounds.

The \_\_\_\_\_\_\_\_\_\_ of a compound tells us;

* Which \_\_\_\_\_\_\_\_\_\_\_\_ are in the compound,
* How \_\_\_\_\_\_\_ atoms of each element there are in the compound.

|  |  |  |
| --- | --- | --- |
| Name of compound | Formula | Atoms in the compound |
| Carbon Dioxide |  | **1** carbon atom**2** oxygen atoms |
| Water |  |  |
| Ammonia |  |  |
| Calcium oxide |  |  |
| Copper Sulphate |  |  |
| Calcium Hydroxide |  |  |

**e.g. Limestone is mainly - Calcium carbonate, CaCO3**

|  |  |  |  |
| --- | --- | --- | --- |
| **Symbol** |  |  |  |
| **Element** |  |  |  |
| **Number of atoms** |  |  |  |

**Describing elements and compounds**

Use the letters in the boxes below to answer the questions that follow.

|  |  |  |
| --- | --- | --- |
| A all the atoms are the same | p314B  | Ccontains more than one kind of atom joined together |
| p314D | Esmallest particle of an element | p314F |
| Gsmall group of atoms held together by bonds | p314H | Iexamples are found in the Periodic Table |

**1** Which box or boxes describe or show a pure element? \_\_\_\_\_\_\_\_\_\_\_\_

**2** Which box or boxes describe or show a pure compound? \_\_\_\_\_\_\_\_\_\_\_

**3** Which box or boxes describe or show a molecule? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**4** Which box shows molecules that contain three atoms? \_\_\_\_\_\_\_\_\_\_\_\_

**5** Which box shows a diagram of an element that is not made up of molecules? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**6** Oxygen gas has the formula O2.Which box shows a diagram of molecules of oxygen gas? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

The drawings below show the particles of atoms and molecules in different substances.

|  |  |  |
| --- | --- | --- |
| p332A | B | C |
| D | E | F |

**1** Which of the boxes shows a pure substance? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**2** Which of the boxes shows an impure substance? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**3** Which of the boxes shows a pure compound? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**4** Which of the boxes shows a pure element? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**5** Match each of the descriptions below with the correct box:

**a** the atoms of solid pure gold (Au) \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**b** the molecules of water (H2O) and carbon monoxide (CO) mixed together \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**c** atoms of the element argon (Ar) as a gas \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**d** a mixture of solid silver (Ag) with a few atoms of copper (Cu) \_\_\_\_\_

**e** pure carbon dioxide (CO2) gas \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**f** a mixture of neon (Ne) and oxygen (O2) gases. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

The air is a **mixture** made up of different elements and compounds. The main gases in air are given in the table below, along with the amount of each found in 100cm3 of dry air. Dry air has no water vapour in it.

**Gases in the air**

|  |  |
| --- | --- |
| **Name of gas** | **Approximate volume of each gas in 100cm3 of air** |
| Nitrogen N2 | 78 |
| Oxygen O2 | 21 |
| Argon Ar | 1 |
| carbon dioxide CO2 | 0.03 |

1. Why would it be so difficult to show the amount of carbon dioxide in the air on a graph? ……………………………………………………………………………………………………
2. Which is the most abundant gas in air? ………………………………………….
3. Which gas makes up about one-fifth of the air? …………………………………….
4. Label each molecule below.



*Atoms and molecules in air*

**1** Air is a *mixture* of different gases.

**a** Colour in the atoms. Use a different colour for each different type of atom.

**b** Fill in the missing parts of the table.

|  |  |  |  |
| --- | --- | --- | --- |
|  | **Name of gas** | **Atoms or molecules?** | **Element or compound?** |
| p326a |  | molecules | element |
|  |  |  |  |
|  |  |  | compound |
|  |  |  |  |
|  | argon |  | element |

**2**

|  |  |
| --- | --- |
| A |  |
| B |  |
| C |  |
| D |  |

This question is about the particles in different substances.

**a** Match up the following descriptions with the correct drawing.

The first one has been done for you.

**i** pure argon

**ii** a mixture of nitrogen and argon

**iii** impure oxygen

**iv** pure water

**b** Colour in the atoms the same as for question **1**.

**Chemical Formula**

A chemical formula tells you the number and type of each atom in a molecule of an element or a compound. Look at each formula on the left and draw a line to the model that best represents it.

|  |  |  |
| --- | --- | --- |
| Ammonia**NH3** |  |  |
| Hydrogen peroxide**H2O2** |  |  |
| Hydrogen chloride**HCl** |  |  |
| Hydrogen sulphide**H2S** |  |  |
| Oxygen**O2** |  |  |
| Methane**CH4** |  |  |

[6]

**C1.1 Atomic structure**

**Electronic structure and configuration of atoms 12/09/18**

An atom can be described as a central positively charged nucleus containing protons and neutrons surrounded by orbiting electrons in shells. The first shell around the nucleus can hold up to 2 electrons. The 2nd can hold 8 and so can the 3rd (**2,8,8,2**).



|  |  |  |  |
| --- | --- | --- | --- |
| Name | Relative charge | Relative mass | Location |
| **Proton** | +1 Positive | 1 | Nucleus |
| **Neutron** | 0 Neutral | 1 | Nucleus |
| **Electron** | -1 Negative | 0 (1/2000) | Outside nucleus |
|  |  |  |  |

**Atomic number and Mass number**

The atomic (proton) number is defined as the number of protons in an atom.

As atoms are neutral (have no overall electrical charge), the number of protons is equal to the number of electrons.

The mass number is defined as the total number of protons plus neutrons in an atom (the heavy bits).

Using the Periodic table, we can calculate the number of Protons, Electrons and Neutrons for all atoms;

Number of neutrons = Mass number – Atomic number