**Chemistry Unit C1: Structures, Trends, Chemical Reactions, Quantitative Chemistry and Analysis**

**C1.9 Chemical Analysis**

|  |  |  |  |
| --- | --- | --- | --- |
| **Content - CCEA Double Award Chemistry 1 – Fort Hill Integrated College** | Got it | Nearly | Haven’t a clue |
| **C1.9 Chemical Analysis** |
| **Assessing purity and separating mixtures** |  |  |  |
| Can you recall that a pure substance is a single element or compound not mixed with any other substance; |  |  |  |
| Can you demonstrate recall that pure elements and compounds melt and boil at specific temperatures and melting point and boiling point can be used to distinguish pure substances from mixtures; |  |  |  |
| Can you demonstrate knowledge and understanding that a formulation is a mixture that has been designed as a useful product and is formed by mixing together several different substances in carefully measured quantities to ensure the product has the required properties, for example alloys, medicines and fertilisers; |  |  |  |
| Can you identify and define the terms soluble, insoluble, solute, solvent, solution, residue, filtrate, distillate, miscible, immiscible, evaporation and condensation; |  |  |  |
| investigate practically how mixtures can be separated using filtration, crystallisation, paper chromatography, simple distillation or fractional distillation (including using fractional distillation in the laboratory to separate miscible liquids, for example ethanol and water); |  |  |  |
| Can you describe paper chromatography as the separation of mixtures of soluble substances by running a solvent (mobile phase) through the mixture on the paper (stationary phase), which causes the substances to move at different rates over the paper; andCan you interpret a paper chromatogram **including calculating Rf values**. |  |  |  |
| Can you analyse given data on mixtures to make judgements on the most effective methods of separation and plan experiments to carry out this separation; |  |  |  |
| Can you describe how to use anhydrous copper(II) sulfate to test for water; |  |  |  |
| **Tests for ions** |  |  |  |
| Can you describe how to carry out a flame test using nichrome wire and concentrated hydrochloric acid to identify metal ions; |  |  |  |
| Can you recall the flame colours of different metal ions:* lithium (crimson);
* sodium (yellow/orange);
* potassium (lilac);
* calcium (brick red); and
* copper(II) (blue–green/green–blue); and
 |  |  |  |
| Can you describe how to *identify the ions in an ionic compound using flame tests**(****Prescribed Practical C2****).* |  |  |  |

**Assessing purity and separating mixtures**

Chemical analysis is extremely important. It is used in forensics, the food industry, Water Board to identify and test the purity of different substances.

**Pure substances**

In chemistry a **pure substance is a single element or compound not mixed with any other substance.**

|  |  |
| --- | --- |
| **Pure substance** | **Mixture** |
|  (only the element carbon) | (O2, N2, CO2, + trace gases) |
| (only H2O) | (water + dissolved salts) |
| (only NaCl) | (water, lactose, fat + minerals) |

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Diamond | Water | Table salt | Air | Mineral water | milk |

**Melting and boiling point of pure substances and mixtures**



**Melting point** is the temperature at which a solid changes into a liquid.

**Boiling point** is the temperature at which a liquid changes into a gas.

**Pure** elements and compounds melt and boil at specific temperatures. Melting point and boiling point can be used to distinguish pure substances from mixtures. These are physical tests, not chemical tests.

e.g. water - boiling point 100ºC; melting point 0 ºC

 sulphur - boiling point 444ºC; melting point 115 ºC



**Impure** substances **do not have sharp** melting or boiling points, but melt/boil over a range of temperatures. The presence of an impurity usually **lowers** the melting point of a solid or raises the boiling point of a liquid.

**Formulations**

A **formulation** is a mixture that has been designed as a useful product and is formed by mixing together several different substances in carefully measured quantities to ensure the product has the required properties. E.g. tomato ketchup, alloys, medicines and fertilisers.



1. The table shows the melting and boiling points of some metallic elements and alloys named by the letters A-C.

|  |  |  |  |
| --- | --- | --- | --- |
|  | Melting point / ºC | Boiling point / ºC | Element or Alloy? |
| **A** | -34 | 356 |  |
| **B** | 420 | 913 |  |
| **C** | 1425 - 1540 | 2530 - 2545 |  |

1. In what state is substance **A** at room temperature and pressure? …………
2. In what state is substance **B** at room temperature and pressure? …………
3. In what state is substance **C** at room temperature and pressure? …………
4. What is an alloy? ……………………………………………………………………………………………………………………
5. Classify the substances in the table as elements or alloys. Explain your answer.

………………………………………………………………………………………………………………………

**Separating mixtures**

A **mixture** is defined as **two or more substances mixed together**. The substances in a mixture are easy to separate as they are not chemically bonded together.

**Key terms**

|  |  |
| --- | --- |
| Term | Definition |
| soluble  | A substance that will dissolve in a solvent |
| insoluble | A substance that will not dissolve in a solvent |
| solute | The substance that dissolves in a solvent |
| solvent | The liquid in which a solute dissolves |
| solution | A solute dissolved in a solvent |
| residue | The solid that remains on the filter paper after filtration |
| filtrate | The liquid which passes through the filter paper during filtration |
| distillate | The liquid that is cooled from the vapour and collected during distillation |
| miscible | Liquids that mix together e.g. alcohol and water |
| immiscible | Liquids which do not mix together but form two distinct layers e.g. oil and water |
| evaporation | The change of state from liquid to gas when heated |
| condensation | The change of state from gas to liquid when cooled |

**Separating techniques**

The method used to separate a mixture depends on the physical properties of the components of the mixture.

1. **Filtration** – used to separate an insoluble solid from a liquid e.g. sand and water



1. **Evaporation** – to separate a dissolved solid from a solvent e.g. salt from a salt solution.
	1. **Crystallisation** (another method of separating a dissolved solid from a solvent)

A solution is heated to boil off some solvent. This creates a **saturated** solution – **one in which no more solid can dissolve at that temperature**. The saturated solution then cools. The solute becomes less soluble and so cannot remain dissolved and crystallises out of solution.

1. **Simple distillation** – used to separate a solvent from a solution

Distillation is evaporation followed condensation (**anti-bumping granules** are added to the mixture in the flask to promote smooth boiling).

1. **Fractional distillation** – used to separate miscible liquids which have different boiling points.

The long **fractionating column** provides better separation of liquids as any evaporated liquids below their boiling point condense in the column and fall back down into the flask. This means the fractions are pure.

***Quick questions***

1. A mixture of insoluble copper (II) carbonate and water was separated by filtration.
	1. Name the filtrate: ……………………………………………………………
	2. Name the residue: ……………………………………………………………
2. Pure water can be obtained by distillation of copper (II) sulfate solution.
	1. Name the solute in copper (II) sulfate solution.

………………………………………………………………………………………………

* 1. Name the change of state which occurs in the distillation flask.

………………………………………………………………………………………………

* 1. What is the distillate?

………………………………………………………………………………………………

* 1. What is added to the distillation flask to ensure smooth boiling? …………………………………………………………
1. **Paper chromatography** – used to separate mixtures of soluble substances in a solution e.g. inks, dyes, food colouring.



**Method**

1. Draw a pencil line 1-2cm from the bottom of chromatography paper.
2. Place a concentrated spot of substance on the pencil line.
3. Place the paper in a beaker with the solvent at the bottom.
4. Leave the solvent to travel up the paper.
5. When the solvent is near the top, take the paper out of the solvent and mark the level that it reached – the **solvent front** (the furthest distance travelled by the solvent)
6. Leave the paper to dry. The mixture should have separated into different components, seen as spots on the paper.

In paper chromatography, the **stationary phase** is the paper and the **mobile phase** is the solvent.

Each substance in the mixture moves at a different rate up the paper (depending on their relative attraction to the paper and the solvent).

**Rf values**

The separated components of the mixtures can be identified by calculating the **Rf** value using the equation:

|  |  |  |
| --- | --- | --- |
| **Rf** | **=** | **distance moved by substance** |
| **distance moved by solvent** |

The **Rf** value of a particular substance in the same solvent is always the same.



Try Q5 C1 H Feb 2019

 **Chemical Test for water**

Water turns **white** anhydrous copper(II) sulphate **blue**.

Anhydrous copper (II) sulfate + water 🡪 hydrated copper (II) sulfate

CuSO4 + 5H2O 🡪 CuSO4.5H2O

***Prescribed Practical C2***

**Cation tests** – cations can be identified by carrying out a flame test;

Method

* Dip a loop of **nichrome wire** into **concentrated hydrochloric acid** and then into the salt to be tested.
* Place the loop into a **blue** Bunsen burner flame and record the first colour observed

|  |  |
| --- | --- |
| **Metal ion present** | **Flame colour** |
| Lithium (Li+) | Crimson |
| Sodium (Na+) | Yellow/orange |
| Potassium (K+) | Lilac |
| Calcium (Ca2+) | Brick-red |
| Copper (Cu2+) | Blue-green/green-blue |



|  |  |  |
| --- | --- | --- |
| Calcium burns **red** and Copper **blue-green**,You are the sweetestThat I’ve ever seen. | Lithium is **crimson**,Sodium is **yella**,You should love meI’m a hell of a fella. | Potassium is **lilac**I’m delighted to sayIf you don’t love meI’ll try being single. |

DAS textbook P234/5 – Practice questions